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Periodic Table With Chemistry Formulas SparkCharts



Periodic Table with Chemistry Formulas

CHEMISTRY FORMULAS

ATOMS AND SUBATOMIC PARTICLES

- A. Atom:** Smallest unit of an element that retains the properties of that element.
- B. Proton:** Positive subatomic particle that is the number of protons is the atomic number of an element. It is the mass of one proton.
- C. Neutron:** Subatomic particle that is the number of neutrons in an atom. It has a mass of 1.
- D. Electron:** Negatively charged subatomic particle that has a mass of 1/1837 that of a proton.

- E. Isotopes:** Atoms of the same element that have different numbers of neutrons. For example, $^{127}_{53}\text{I}$ and $^{131}_{53}\text{I}$ are isotopes of iodine.
- F. Relative atomic mass:** The number of protons and neutrons in an atom.
- G. Avogadro's number:** 6.022×10^{23} atoms/mole.

- H. Electron shells:** The number of electrons that can be in each shell of an atom.

Atomic	Mass	Mass No.	Symbol	Name
Proton	1.6726219 × 10 ⁻²⁴	1	p^+	Hydrogen
Neutron	1.6749274 × 10 ⁻²⁴	1	n^0	
Electron	9.1093895 × 10 ⁻³¹	1	e^-	

MOLECULAR MASS

- A. Average molecular mass:** $M_{\text{avg}} = \frac{m_1 \times n_1 + m_2 \times n_2 + \dots + m_n \times n_n}{n_1 + n_2 + \dots + n_n}$
- B. The atomic mass of an element equals the mass of one mole of that element.**
- C. The mass of one mole of atoms equals the mass of one mole of that element.**

- D. The mass of one mole of atoms equals the mass of one mole of that element.**

MOLECULAR FORMULA AND EMPirical FORMULA

- A. Molecular formula:** The chemical formula of a compound that gives the actual number of each kind of atoms in a molecule.
- B. The empirical formula:** The formula C_2H_4 .
- C. Empirical formula:** A formula that contains the lowest whole-number ratio of atoms.

- D. The empirical formula for glucose is $\text{C}_6\text{H}_{12}\text{O}_6$.**

BALANCE CHEMICAL EQUATIONS

- A. Check for charge:** Add the numbers of H^+ and O^- .
- B. Balance metals.**
- C. Balance oxygen.**
- D. Balance hydrogen.**
- E. Use trial and error:** Sometimes, it is necessary to multiply the entire equation by a number to get the correct coefficients.
- F. Rewrite the equation so that there are no fractions.**
- G. Check for charge in the reaction products.**

PERCENTAGE AND RATE

- A. Percentage composition:** Total mass of each element in a compound divided by the total mass of the compound.
- B. Part per million:** Expresses the number of grams of a

PROBLEMS INVOLVING MOLES

1. CONCENTRATION (MOLARITY)

A. Definition: Molarity = Number of moles of solute per liter of solution.

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

B. Units of molarity:

$$\text{Molarity} = \frac{\text{moles}}{\text{liters}}$$

C. Molar mass:

$$\text{Molar mass} = \frac{\text{mass}}{\text{moles}}$$

D. Molar concentration:

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Synopsis

SparkCharts™: The information you need-concisely, conveniently, and accurately. ® Created by Harvard students for students everywhere, these study companions and reference tools cover a wide range of college and graduate school subjects, from Business and Computer Programming to Medicine, Law, and Languages. They'll give you what it takes to find success in school and beyond. Outlines and summaries cover key points, while diagrams and tables make difficult concepts easier to grasp.® This two-page chart is a perfect reference for homework and problem sets. On side one, the chart lists the most important chemical formulas and provides quick refreshers on significant figures and balancing equations. Side two includes a beautiful periodic table that gives the following information for each element: Name, Atomic number, Atomic symbol, Atomic mass, Oxidation states (most stable state in bold), Electronegativity, Electron affinity, First ionization potential, Atomic radius, and Electron shell configuration.

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Very useful for my chem class at the university of Minnesota

A nice-looking periodic table for quick reference! What more could you want?

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