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Periodic Table With Chemistry Formulas SparkCharts

* Periodic Table with Chemistry Formulas		
CHEMISTRY FORMULAS		
ATOMS AND SUBATOMIC PARTICLES 1. Atomic number: The number of protons in an atom of an element, represented by the symbol Z . No. of e^- = atomic number 2. Mass number: The number of protons plus the number of neutrons in an atom of an element, represented by the symbol A . No. of p^+ + No. of n^0 = mass number 3. Isotopes: Different versions of an element that possess different numbers of neutrons, and therefore different mass numbers. No. Neutrons = $A - Z$ (Atomic No.) 4. Calculating atomic mass: The atomic mass (mass number) of an element is the weighted average of the masses of its isotopes. 5. Atomic mass: The mass of an atom of an element, represented by the symbol A_r . Atomic mass = $\frac{\text{mass of atom}}{\text{mass of } ^{12}\text{C}}$ 6. Relative atomic mass: The relative atomic mass of an element is the weighted average of the masses of its isotopes. Relative atomic mass = $\frac{\text{mass of atom}}{\text{mass of } ^{12}\text{C}}$ 7. Molar mass: The mass of one mole of a substance, represented by the symbol M_r . Molar mass = $\frac{\text{mass of substance}}{\text{number of moles}}$ 8. Avogadro's number: 6.02×10^{23} particles per mole. 9. The atomic mass of an element is the mass of one atom of the element. 10. The mass of one mole of an element is the mass of the element.	PERCENTAGE COMPOSITION 1. Calculating percentage composition: The percentage composition of a compound is the mass of each element in a compound, expressed as a percentage of the total mass. Percentage composition = $\frac{\text{mass of element}}{\text{total mass}} \times 100$ 2. Empirical formula: The simplest formula of a compound, representing the relative numbers of atoms of each element in a compound. 3. Molecular formula: The actual formula of a compound, representing the actual numbers of atoms of each element in a molecule. 4. Relative molecular mass: The relative molecular mass of a compound is the mass of one molecule of the compound, expressed as a multiple of the mass of one atom of ^{12}C . Relative molecular mass = $\frac{\text{mass of molecule}}{\text{mass of } ^{12}\text{C}}$ 5. Relative atomic mass: The relative atomic mass of an element is the mass of one atom of the element, expressed as a multiple of the mass of one atom of ^{12}C . Relative atomic mass = $\frac{\text{mass of atom}}{\text{mass of } ^{12}\text{C}}$ 6. Molar mass: The mass of one mole of a substance, represented by the symbol M_r . Molar mass = $\frac{\text{mass of substance}}{\text{number of moles}}$ 7. Avogadro's number: 6.02×10^{23} particles per mole. 8. The atomic mass of an element is the mass of one atom of the element. 9. The mass of one mole of an element is the mass of the element.	PERCENTAGE COMPOSITION 1. Calculating percentage composition: The percentage composition of a compound is the mass of each element in a compound, expressed as a percentage of the total mass. Percentage composition = $\frac{\text{mass of element}}{\text{total mass}} \times 100$ 2. Empirical formula: The simplest formula of a compound, representing the relative numbers of atoms of each element in a compound. 3. Molecular formula: The actual formula of a compound, representing the actual numbers of atoms of each element in a molecule. 4. Relative molecular mass: The relative molecular mass of a compound is the mass of one molecule of the compound, expressed as a multiple of the mass of one atom of ^{12}C . Relative molecular mass = $\frac{\text{mass of molecule}}{\text{mass of } ^{12}\text{C}}$ 5. Relative atomic mass: The relative atomic mass of an element is the mass of one atom of the element, expressed as a multiple of the mass of one atom of ^{12}C . Relative atomic mass = $\frac{\text{mass of atom}}{\text{mass of } ^{12}\text{C}}$ 6. Molar mass: The mass of one mole of a substance, represented by the symbol M_r . Molar mass = $\frac{\text{mass of substance}}{\text{number of moles}}$ 7. Avogadro's number: 6.02×10^{23} particles per mole. 8. The atomic mass of an element is the mass of one atom of the element. 9. The mass of one mole of an element is the mass of the element.



Synopsis

SparkCharts®: The information you need—concisely, conveniently, and accurately. Created by Harvard students for students everywhere, these study companions and reference tools cover a wide range of college and graduate school subjects, from Business and Computer Programming to Medicine, Law, and Languages. They'll give you what it takes to find success in school and beyond. Outlines and summaries cover key points, while diagrams and tables make difficult concepts easier to grasp. This two-page chart is a perfect reference for homework and problem sets. On side one, the chart lists the most important chemical formulas and provides quick refreshers on significant figures and balancing equations. Side two includes a beautiful periodic table that gives the following information for each element: Name Atomic number Atomic symbol Atomic mass Oxidation states (most stable state in bold) Electronegativity Electron affinity First ionization potential Atomic radius Electron shell configuration

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Customer Reviews

Very useful for my chem class at the university of Minnesota

A nice-looking periodic table for quick reference! What more could you want?

Item arrived promptly and as described! Would use again!

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